

**(WO/2003/011233) COSMETIC RAW MATERIAL HAVING IMPROVED PROPERTIES
AND PROCESSES FOR PREPARING THE SAME**

Biblio. Data Description Claims National Phase Notices Documents

Latest bibliographic data on file with the International Bureau

Pub. No.: WO/2003/011233 International Application No.: PCT/KR2002/001410
 Publication Date: 13.02.2003 International Filing Date: 26.07.2002

IPC: A61K 8/25 (2006.01), A61K 8/26 (2006.01), A61K 8/27 (2006.01), A61K 8/365 (2006.01), A61K 8/368 (2006.01), A61K 8/49 (2006.01), A61K 8/67 (2006.01), A61Q 19/00 (2006.01), A61Q 19/02 (2006.01), A61Q 19/08 (2006.01)

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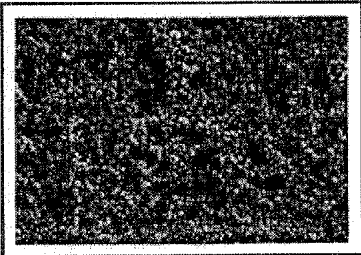
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Priority Data: 2001/45703 28.07.2001 KR
 2001/45697 28.07.2001 KR
 2001/48177 10.08.2001 KR

Title: COSMETIC RAW MATERIAL HAVING IMPROVED PROPERTIES AND PROCESSES FOR PREPARING THE SAME

Abstract: The present invention provides a hybrid material comprised of an active component for raw materials for cosmetics and a layered metal hydroxide, showing good stability, low toxicity, low irritation, good sustained release and good dispersibility. The present invention provides a method for preparing the hybrid material, using the coprecipitation method, the ion-exchange method or the adsorption method, depending on properties of the active component for raw materials for cosmetics. The method may further comprise a step of coating the surface of the hybrid material after preparation of the hybrid material. The present invention provides cosmetics comprising the above hybrid material.



Designated States: AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZM, ZW.
 African Regional Intellectual Property Org. (ARIPO) (GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW)
 Eurasian Patent Organization (EAPO) (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM)
 European Patent Office (EPO) (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, SK, TR)
 African Intellectual Property Organization (OAPI) (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

Publication Language: English (EN)
 Filing Language: Korean (KO)

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
13 February 2003 (13.02.2003)

PCT

(10) International Publication Number
WO 03/011233 A1

(51) International Patent Classification⁷: **A61K 7/00**, 7/48

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(21) International Application Number: PCT/KR02/01410

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(22) International Filing Date: 26 July 2002 (26.07.2002)

(25) Filing Language: Korean

(26) Publication Language: English

(30) Priority Data:
2001/45703 28 July 2001 (28.07.2001) KR
2001/45697 28 July 2001 (28.07.2001) KR
2001/48177 10 August 2001 (10.08.2001) KR

(81) Designated States (*national*): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZM, ZW.

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(84) Designated States (*regional*): ARIPO patent (GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, SK, TR), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

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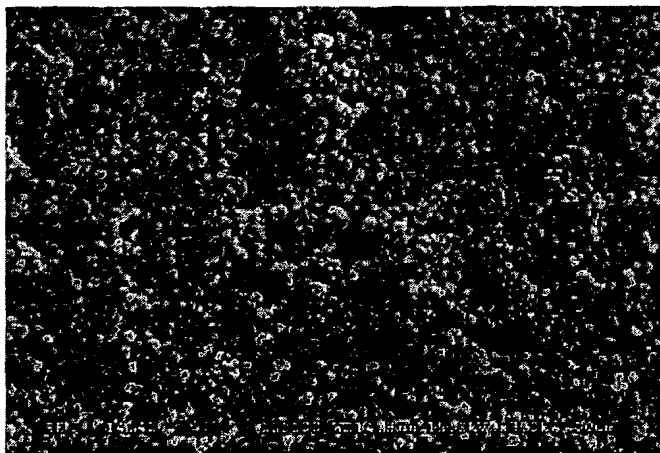
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Published:

— with international search report

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: COSMETIC RAW MATERIAL HAVING IMPROVED PROPERTIES AND PROCESSES FOR PREPARING THE SAME



(57) Abstract: The present invention provides a hybrid material comprised of an active component for raw materials for cosmetics and a layered metal hydroxide, showing good stability, low toxicity, low irritation, good sustained release and good dispersibility. The present invention provides a method for preparing the hybrid material, using the coprecipitation method, the ion-exchange method or the adsorption method, depending on properties of the active component for raw materials for cosmetics. The method may further comprise a step of coating the surface of the hybrid material after preparation of the hybrid material. The present invention provides cosmetics comprising the above hybrid material.



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COSMETIC RAW MATERIALS HAVING IMPROVED PROPERTIES AND PROCESSES FOR PREPARING THE SAME

TECHNICAL FIELD

5 The present invention relates to a hybrid material for use as raw materials for cosmetics and a method of preparation thereof. More specifically, it relates to a hybrid material for raw materials for cosmetics showing good safety, low toxicity, low irritation, good sustained release, good dispersibility, etc., and a method of preparation thereof.

BACKGROUND ART

10 Human skin includes two layers, the epidermis and dermis. The epidermis consists of Keratinocytes which are keratin-forming cells, Melanocytes which are melanin-containing cells, and Langerhans which are cells related to immune function
15 in the epidermis. The dermis has collagen which is an extra-cellular matrix (ECM) component, and fibroblast forming glycosaminoglycans (GAGs). These components must maintain their functions to keep human skin healthy and beautiful. However, as humans get older or their skin exposed to harmful environments such as ultraviolet rays, the concentration of reactive oxygen species (ROS) increases to
20 accelerate oxidation, thereby resulting in phenomena such as pigmentary deposit, keratinization and wrinkles.

 Thus, there have been various efforts to prevent or inhibit natural aging or temporary skin troubles occurring in skin. As part thereof, various natural and synthetic materials have been developed and used as raw materials for cosmetics for
25 skin care. Such raw materials for cosmetics may be classified in terms of their functions into anti-oxidant, keratin remover, skin whitening agent, spot/freckle remover, skin calmativ, skin moisturizing agent, skin activator, free blight preventing agent, skin metabolism promoter, wrinkle preventing agent, etc.

 Vitamins can be mentioned as one of most representative raw materials for
30 cosmetics. Vitamins are essential materials for the body, having functions such as promoting body metabolism, anti-oxidation, protecting cell wall, enhancing immunity, increasing resistance to infection, etc. For human beings, biosynthesis of vitamins is impossible and thus intake of vitamins through food is essential. When vitamins are deficient, various diseases occur. In addition, vitamins play an
35 important part in skin care and treatment such as preventing pigmentary deposit, promoting synthesis of collagen, blocking ultraviolet rays, preventing drying and

keratinization of skin, preventing wrinkles and moisturizing skin. Exemplary vitamins include vitamin A such as retinol, vitamin C such as ascorbic acid, vitamin E such as tocopherol, and derivatives thereof.

In addition to vitamins, important raw materials for cosmetics include α -hydroxy acid (AHA) such as lactic acid, citric acid and salicylic acid, which functions to remove keratin layer of skin and thereby accelerate metabolism of skin; kojic acid (5-hydroxy-2-(hydroxymethyl)-4H-pyran-4-on) which functions to whiten skin by inhibiting melanin biosynthesis; indol-3-acetic acid ($C_{10}H_9NO_2$) which functions to prevent wrinkles by accelerating fibroblast multiplication; salicylic acid (2-hydroxy benzoic acid; $C_7H_6O_3$) which functions as an anti-oxidant to treat pimples; and other various known raw materials for cosmetics.

However, most of the aforementioned raw materials for cosmetics are limited in practical use because of problems in their inherent stability, skin irritation and toxicity, sustained release, dispersibility, etc., and thus have not exerted their functions and effects sufficiently. For example, vitamins are very unstable physicochemically and easily destroyed by heat, light, moisture, oxygen, alkali, etc., leading to deterioration of their functions and effects, or discoloration or malodor. Also, α -hydroxy acid such as lactic acid irritates skin. Kojic acid degenerates melanin pigments and melanin cells by penetrating into skin base layer, and thus if used in high concentration, induces various skin diseases such as dermatitis and skin cancer. It is also oxidized under light and high temperature to cause discoloration. Furthermore, indol-3-acetic acid is unstable under external environment, heat, light, moisture, oxygen, etc., especially light, to cause discoloration and malodor.

There have been many studies on formulation technologies to stabilize raw materials for cosmetics or reduce skin irritation or toxicity. In particular, such studies have been performed for vitamins, and based on them, studies on materials other than vitamins have been done.

Major prior stabilization methods for vitamins include oil-in-water (O/W) or water-in-oil (W/O) emulsion technology, multiple liquid crystal film technology, liposome technology, porous bead technology, water-soluble polymer technology and polysaccharide capsule technology.

Korean Patent No. 115706 discloses a method for preparing a vitamin nano-capsule by impregnating and sealing vitamin and various active components within an inner phase fine globule, and further stabilizing this using a double-layer lipid film. Furthermore, Korean Patent Laid-open Publication No. 2000-0048451 discloses a method for capsuling a lipophilic active component globule core with

water insoluble synthetic or natural water-dispersible anionic polymers. Korean Patent Laid-open Publication No. 2000-0069893 discloses a method for stabilization by impregnating an active component in polysaccharide globule such as starch followed by coating the exterior with water-soluble polymers, or a method for stabilization by micro-emulsifying the active component and stabilizing the exterior using an anionic, non-ionic, cationic or amphoteric surfactant. Korean Patent No. 202401 discloses a method wherein an active component is impregnated in a polysaccharide and the exterior is capsuled with a neutral liquid crystal. Korean Patent No. 236484 discloses a method wherein a stable emulsion containing a water-soluble active component is formed and the exterior is further stabilized with a gelatin capsule. Korean Patent Laid-open Publication No. 1999-0070885 discloses a method wherein a mixture of vitamins such as water-soluble vitamins and fat-soluble vitamins is mixed with solid powder such as mica thereby absorbing into the powder and the surface thereof is coated with highly viscous liquid oil. Korean Patent No. 222000 discloses a stabilization method by applying simultaneously a liposome technology and a technology for forming liquid crystal. Korean Patent Laid-open Publication No. 2001-0002411 discloses a stabilization method by combination with natural materials such as coffee beans, tea leaves and cacao extracts.

Methods of using more stable derivative have been reported, but it has some problems. For example, magnesium ascorbyl phosphate can be easily converted into a form of biocompatible L-ascorbic acid compared to other derivatives, but it is difficult to be absorbed into skin. Ascorbyl palmitate is easily absorbed into skin, but difficult to be converted into a form of L-ascorbic acid.

Technologies for stabilizing vitamin A include oil-in-water or water-in-oil emulsion formulation technologies using an oil-soluble anti-oxidant such as butylated hydroxytoluene (BHT) or butylated hydroxy anisole (BHA), a water-soluble anti-oxidant such as ascorbic acid, or a chelating agent such as α -tocopherol. See Korean Patent Laid-open Publication Nos. 1990-021511 and 1999-0018726; US Patent Nos. 4,247,547; 4,826,828; 4,720,353; and 4,466,805; and EP 0 440 398 B1.

Japanese Patent Laid-open Publication No. Hei 2-83309 discloses a technology wherein a water-soluble L-ascorbic acid derivative is introduced into the inner phase of a water-in-oil emulsion to protect from oxidation or decomposition. Japanese Patent Laid-open Publication No. Sho 55-64511 discloses a technology wherein higher fatty acid esters are included within cyclodextrine. Japanese Patent Laid-open Publication No. Hei 3-5326 discloses a technology wherein L-ascorbic

acid phosphoric acid esters or sulfuric acid esters are added to liposome to improve storage stability. Japanese Patent Laid-open Publication No. Hei 5-345714 discloses a technology wherein L-ascorbic acid esters are combined with certain water-soluble polymers to improve stability. All of these patent publications
5 disclose using ester derivatives of L-ascorbic acid, which improves stability compared to L-ascorbic acid itself, but are known to be less effective in reducing melanin or inhibiting tyrosinase activities.

Korean Patent Laid-open Publication No. 2000-0026158 discloses a stabilization method using an oil-in-water emulsion which encapsulate retinol or
10 derivatives based on a matrix of collagen and chitosan. However, retinol or derivatives are known to be thermodynamically very unstable. In particular, they are easily oxidized by oxygen in air, light and heat, resulting in loss of their effects and causing irritation to skin by decomposed retinol derivatives. BHA or BHT of the prior art may be added to stabilize retinol derivatives and reduce discoloration,
15 but these components are also irritant to skin and recently avoided. There have been recent developments proposed on a number of methods for encapsulating with gelatin, alginic acid, etc., but these methods cause excess retinol to burst out at once and irritate skin.

The following methods are known. A method of microcapsuling
20 tocopherol contained in the inner phase of retinyl palmitate, a fat-soluble active component with film material of crosslinked collagen and glucoseaminogruane (Korean Patent Laid-open Publication No. 1999-0065437); a method of using double-capsuling consisting of first encapsulating an inner phase of retinol and secondly capsuling with gellane gum (Korean Patent Laid-open Publication No.
25 1998-21511); a method of capsuling in liposome (Korean Patent Laid-open Publication No. 1998-703668); a method of stabilizing retinyl palmitate in liquid crystal gel phase (Korean Patent Laid-open Publication No. 1996-033439). Additionally, a number of multi-emulsified forms, particularly W/O/W were also proposed to improve shortcomings of the prior single-emulsified forms such as W/O
30 or O/W: see, US Patent No. 3399263, Great Britain Patent No. 1235667, Japanese Patent Laid-open Publication No. Sho 52-134029, Japanese Patent Laid-open Publication No. Sho 53-31578, Japanese Patent Laid-open Publication No. Sho 57-15829, Japanese Patent Laid-open Publication No. Sho 58-183611, Japanese Patent Laid-open Publication No. Sho 59-80326, Japanese Patent Laid-open Publication No.
35 Sho 59-127646, Japanese Patent Laid-open Publication No. Sho 62-2561, Japanese Patent Laid-open Publication No. Sho 63-33311, Korean Patent Laid-open

Publication No. 1995-0000131. There were also known methods wherein an active component is included in pores of porous microbeads in a similar manner to microcapsule (US Patent No. 5,145,675); and a method wherein the surface of porous microbeads containing an active component is treated with silicone or derivatives thereof (Korean Patent Laid-open Publication No. 2000-0067126).

As described above, the prior methods for stabilizing vitamins basically comprises combining a water-soluble or fat-soluble vitamin with oily or water phase polymers, surfactants and other organics thereby forming an inner phase, and if necessary, further encapsulating the exterior with polymers, surfactants, polysaccharides, oils, etc., although slightly different in specific matters.

In the case of α -hydroxy acid (AHA) such as lactic acid, citric acid and salicylic acid, there is no stability problem in using them as active components as raw materials for cosmetics, but if released in excess and applied to skin, they cause severe skin irritation such as torrid heat feeling and prickling. Thus, the amounts must be controlled when used as raw materials for cosmetics.

For kojic acid, Korean Patent Laid-open Publication No. 1998-044041 discloses a technology for preventing low molecular weight kojic acid from penetrating into skin, wherein a kojic acid is converted to a higher molecular weight by e.g. polymerizing the kojic acid using a kojic acid/amino acid complex. Furthermore, Korean Patent Laid-open Publication No. 2002-0025356 discloses a method for solving the problems of potency reduction and browning due to oxidation of kojic acid, using a kojic acid/amino acid complex and concentrated green tea extracts having the anti-oxidation effects. US Patent No. 4847074 discloses a technology for preventing browning and improving stability by including kojic acid within cyclodextrine. Furthermore, US Patent No. 4369174 discloses a method for improving stability against heat and light using derivatives obtained from the esterification reaction of kojic acid and aliphatic carboxylic acid.

In the case of indol-3-acetic acid, since its amount in a formulation is extremely small e.g. about 300 ppm, almost no separate report has been made on a stabilization method. Thus, prior stabilization methods for vitamins are mostly used for it. One prior art for indol-3-acetic acid is Korean Patent Application No. 2001-45629 which discloses a method for using it as a raw material for cosmetics by stabilizing it in the form of liposome.

The prior arts as described above have their own characteristics in degree of stabilization of each active component, release, delivery, biocompatibility, convenience, cost, and the like. However, there still remains a need for further

improvements in safety, stability, sustained release, cost, and the like.

DISCLOSURE OF INVENTION

Therefore, the object of the present invention is to provide raw materials for
5 cosmetics showing good safety, low toxicity, low irritation, good sustained release
and good dispersibility.

Another object of the present invention is to provide a method for preparing
the aforementioned raw materials for cosmetics, economically and effectively.

Another object of the present invention is to provide cosmetics comprising
10 the aforementioned raw materials for cosmetics showing good safety, low toxicity,
low irritation, good sustained release and good dispersibility.

After making efforts to accomplish the above objects, it was discovered that
hybrid materials of active components for raw materials for cosmetics and layered
metal hydroxides show good safety, low toxicity, low irritation, good sustained
15 release and good dispersibility.

BRIEF DESCRIPTION OF DRAWINGS

Fig. 1 is a powder X-ray diffraction pattern for a hybrid material of metal
hydroxide and ascorbate (vitamin C) in the form of hydrozincyte prepared by the
20 coprecipitation method.

Fig. 2 is a scanning electron microgram for a hybrid material of metal
hydroxide and ascorbate (vitamin C) in the form of hydrozincyte prepared by the
coprecipitation method.

Fig. 3 is a scanning electron microgram for a hybrid material of metal
25 hydroxide and ascorbate (vitamin C) in the form of hydrozincyte prepared by the
coprecipitation method using $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$.

Fig. 4 is a powder X-ray diffraction pattern for a hybrid material of metal
hydroxide and retinic acid in the form of hydrozincyte prepared by the
coprecipitation method.

30 Fig. 5 is a scanning electron microgram for a hybrid material of metal
hydroxide and retinic acid in the form of hydrozincyte prepared by the
coprecipitation method.

Fig. 6 is an X-ray diffraction pattern for a hybrid material of $\text{Zn}_3\text{Al-LDH-D-}$
 α -tocopherol succinate and retinol obtained by adsorbing vitamin A (retinol) into the
35 layers of the hybrid material of $\text{Zn}_3\text{Al-LDH}$ and D- α -tocopherol succinate prepared
by the coprecipitation method.

Fig. 7 is a field emission scanning electron microgram for $\text{Zn}_5(\text{OH})_8(\text{C}_3\text{H}_5\text{O}_3)_2 \cdot n\text{H}_2\text{O}$ particles prepared by the coprecipitation method.

Fig. 8 is a powder X-ray diffraction pattern for $\text{Zn}_6\text{Al}_2(\text{OH})_{16}(\text{C}_6\text{H}_5\text{O}_7)_{2/3} \cdot 4\text{H}_2\text{O}$ by the coprecipitation method.

5 Fig. 9 is a powder X-ray diffraction pattern for a hybrid material of metal hydroxide and kojic acid in the form of hydrozincyte prepared by the coprecipitation method.

Fig. 10 is a powder X-ray diffraction pattern for $\text{Zn}_5(\text{OH})_8(\text{C}_7\text{H}_5\text{O}_3)_2 \cdot n\text{H}_2\text{O}$ particles prepared by the coprecipitation method.

10 Fig. 11 is a powder X-ray diffraction pattern for $\text{Zn}_5(\text{OH})_8(\text{C}_{10}\text{H}_8\text{NO}_2)_2 \cdot n\text{H}_2\text{O}$ particles prepared by the coprecipitation method.

Fig. 12 is a field emission scanning electron microgram for $\text{Zn}_5(\text{OH})_8(\text{C}_{10}\text{H}_8\text{NO}_2)_2 \cdot n\text{H}_2\text{O}$ particles prepared by the coprecipitation method.

15 Fig. 13 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, surface-coated with silica using tetraethoxysilicate (TEOS), the hybrid material being in the form of hydrozincyte in slurry phase prepared by the coprecipitation method.

Fig. 14 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, surface-coated with silica by adding
20 tetraethoxysilicate (TEOS) to a suspension of the hybrid material without additional separation or washing step, the hybrid material being in the form of hydrozincyte prepared by the coprecipitation method.

Fig. 15 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, surface-coated with silica using tetraethoxysilicate
25 (TEOS), the hybrid material being in powder form prepared by the coprecipitation method.

Fig. 16 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, surface-coated with silica using tetramethoxysilicate (TMOS), the hybrid material being in the form of hydrozincyte in slurry phase
30 prepared by the coprecipitation method.

Fig. 17 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, surface-coated with silica using tetraethoxysilicate (TEOS) and an acid catalyst, the hybrid material being in the form of hydrozincyte in slurry phase prepared by the coprecipitation method.

35 Fig. 18 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, surface-coated with silica using tetraethoxysilicate

(TEOS), the hybrid material being in the form of ZnAl-LDH in slurry phase prepared by the ion-exchange method.

Fig. 19 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, secondary-coated with silane and
5 polymethylmethacrylate (PMMA), the hybrid material being in the form of hydrozincyte prepared by the coprecipitation method.

Fig. 20 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, secondary-coated with polymethylmethacrylate (PMMA) by solution polymerization method, the hybrid material being in the form
10 of hydrozincyte coated with silica.

Fig. 21 is a scanning electron microgram for particles of a hybrid material of metal hydroxide and vitamin C, secondary-coated with ethyl cellulose in a solution, the hybrid material being in the form of hydrozincyte coated with silica.

Fig. 22 is a scanning electron microgram for particles of a hybrid material of
15 metal hydroxide and vitamin C, secondary-coated with polyvinyl alcohol (PVA) and polymethylmethacrylate (PMMA) in a solution, the hybrid material being in the form of hydrozincyte coated with silica.

Fig. 23 is a scanning electron microgram for a hybrid material of $\text{Zn}_5(\text{OH})_8(\text{C}_{10}\text{H}_8\text{NO}_2)_2 \cdot n\text{H}_2\text{O}$ obtained in Example 11, surface-coated with silica
20 using tetraethoxysilicate (TEOS).

Fig. 24 is a graph showing the storage stability evaluation results of vitamin C hybrid materials prepared in Example 12.

Fig. 25 is a graph showing the time-dependent elution pattern of vitamin C hybrid materials prepared in Examples 3 and 12.

25 Fig. 26 is a graph showing the time-dependent elution pattern of indol-3-acetic acid in 0.8 % NaCl aqueous solution for $\text{Zn}_5(\text{OH})_8(\text{C}_{10}\text{H}_8\text{NO}_2)_2 \cdot n\text{H}_2\text{O}$ hybrid materials prepared by the coprecipitation method.

BEST MODES FOR CARRYING OUT THE INVENTION

30 "Layered metal hydroxides" as used in the present invention are weakly basic or neutral inorganic materials having good skin affinity without adverse effects such as skin irritation. Furthermore, the layered metal hydroxides have layered crystal structure and show peculiar reactivity between the layers. Also, they have anion exchange capacity, and thus can be effectively used for anionic components
35 among active cosmetic components.

Layered metal hydroxides used in the present invention are preferably

hydroxy double salts (HDS) or layered double hydroxides (LDHs). Among hydroxy double salts, those having a crystal structure of hydrozincyte ($\text{Zn}_5(\text{OH})_6(\text{CO}_3)_2 \cdot n\text{H}_2\text{O}$) or zinc basic salts ($\text{Zn}_5(\text{OH})_8\text{Cl}_2 \cdot n\text{H}_2\text{O}$) are more preferable. Layered double hydroxides are compounds such as hydrotalcite having a formula of

5 $[\text{M}(\text{II})_{1-x}\text{M}(\text{III})_x(\text{OH}_2)]^{x+}(\text{A}^{n-}_{x/n}) \cdot m\text{H}_2\text{O}$, where M(II) is a divalent metal cation such as Zn(II), Mg(II), Ca(II) and Fe(II); M(III) is a trivalent metal cation such as Al(III), Fe(III), Cr(III) and Co(III); and A is an anion. All these compounds have positive charges within hydroxide layers by different causes, and thus anions are present between the layers to compensate for the positive charges and may be replaced by

10 other anionic species. Therefore, these compounds have capacity to stabilize negatively charged inorganic ions, organic ions or biomolecules between the layers.

Active components used in raw materials for cosmetics according to the present invention include natural or synthetic organic materials having the functions such as anti-oxidation, removing keratin, skin whitening, removing spots and

15 freckles, skin sedation, skin moisturizing, skin activation, preventing fire blight, accelerating skin metabolism and preventing wrinkles. Examples include vitamins, α -hydroxy acids such as lactic acid, citric acid and salicylic acid, kojic acid (5-hydroxy-2-(hydroxymethyl)-4H-pyran-4-on), indol-3-acetic acid and salicylic acid (2-hydroxy benzoic acid).

20 Vitamins include, by way of example, vitamin A, vitamin C, vitamin E, etc., including pure vitamins and derivatives thereof. Examples of vitamin A include retinic acid ($\text{C}_{20}\text{H}_{28}\text{O}_2$), retinol ($\text{C}_{20}\text{H}_{30}\text{O}$), retinal ($\text{C}_{20}\text{H}_{28}\text{O}$), retinyl acetate ($\text{C}_{22}\text{H}_{32}\text{O}_2$) and retinyl palmitate ($\text{C}_{36}\text{H}_{60}\text{O}_2$). Examples of vitamin C include ascorbic acid ($\text{C}_6\text{H}_8\text{O}_6$) and ascorbate. Examples of vitamin E include α -

25 tocopherol ($\text{C}_{29}\text{H}_{50}\text{O}_2$), tocopherol succinate ($\text{C}_{33}\text{H}_{54}\text{O}_5$) and tocopheryl acetate ($\text{C}_{31}\text{H}_{52}\text{O}_3$).

Furthermore, the present invention provides a method for preparing the aforementioned hybrid materials of active components for raw materials for cosmetics and layered metal hydroxides.

30 One embodiment of the method utilizes coprecipitation. This method comprises dissolving an active component for raw materials for cosmetics and starting materials for a layered metal hydroxide in an aqueous solvent or a mixed solvent, thereby causing precipitation. If necessary, a base may be added to induce precipitation.

35 The starting materials for the layered metal hydroxide are preferably those having high solubilities which are used in solvents and effectively cause precipitation

in the form of layered metal hydroxides when a base is added. Therefore, although the layered metal hydroxide can be used directly, it is more preferable to use metal salts thereof such as metal carbonate, metal chlorides, metal nitrates and organic metal salts (e.g. acetate, oxalate and citrate). Examples include ZnCl_2 , $\text{Zn}(\text{NO}_3)_3$,
5 $\text{Zn}(\text{CH}_3\text{COO})_2$, MgCl_2 , AlCl_3 , CaCl_2 or hydrates thereof.

The coprecipitation method is particularly suitable for active components having negative charges. Examples of such active components include retinic acid, L-ascorbic acid, tocopherol succinate, lactic acid, citric acid, salicylic acid, kojic acid, and indol-3-acetic acid.

10 A concentration of the metal ion in the coprecipitation method is e.g. 0.01 to 5 M. The amount of the active component is e.g. 0.1 to 10 equivalent moles relative to total moles of the metal components. Added bases include for example alkali metal hydroxide or an amine. The reaction solution is kept at pH 4 to 11, preferably pH 6 to 8, and the reaction temperature is 0 to 50 °C, preferably 0 to 10 °C.
15 The reaction time is preferably 30 minutes or more. Furthermore, reaction is preferably performed with continuous purge by nitrogen or other inert gas and under shield.

Another embodiment of the method for preparing hybrid materials of active components for raw materials for cosmetics and layered metal hydroxides according
20 to the present invention utilizes ion-exchange. This method comprises

(1) reacting starting materials for a layered metal hydroxide in a basic solution to form a layered metal hydroxide salt, and

(2) dispersing the layered metal hydroxide salt in an aqueous solvent or a mixed solvent, and dissolving an active component for raw materials for cosmetics,
25 thereby causing ion-exchange reaction.

The ion-exchange method has a shortcoming in that the preparation process becomes more complex compared to the coprecipitation method. However, it is preferable as an alternative in a case where preparing the hybrid materials by the coprecipitation method is unsuitable.

30 Starting materials for layered metal hydroxides used in the ion-exchange method are the same as those described in the coprecipitation method. The amount of the active component added is e.g. 0.1 to 10 equivalent moles relative to the amount of metal hydroxide salts used. The reaction solution in the ion-exchange reaction is kept at e.g. about pH 5 to 11, preferably about pH 7 to 8. The reaction
35 temperature is 0 °C to 50 °C, preferably 0 to 10 °C. The reaction time is preferably 30 minutes or more. Furthermore, the reaction is preferably performed with

continuous purge by nitrogen or other inert gas and under shield.

Another embodiment of the method for preparing hybrid materials of active components for raw materials for cosmetics and layered metal hydroxides according to the present invention utilizes adsorption. This method comprises

5 (1) reacting starting materials for a layered metal hydroxide and an organic acid or an anionic surfactant in an aqueous solvent or a mixed solvent to form organic-inorganic layered hybrid material, and

(2) dispersing the said organic-inorganic layered hybrid material in an aqueous solvent or a mixed solvent, and adding an active component for raw
10 materials for cosmetics, thereby adsorbing the active components onto the organic-inorganic layered hybrid material.

The adsorption method is particularly useful for neutral or highly hydrophobic active components. Such active components include retinol, retinal, retinyl acetate, retinyl palmitate, α -tocopherol and tocopheryl acetate.

15 In the adsorption method, suitable organic materials must be introduced onto the surface of the layered metal hydroxides when preparing the organic-inorganic layered hybrid materials. In order to accomplish this, the organic acid or anionic surfactant is introduced to make the surface of the layered metal hydroxides hydrophobic in the present invention. Examples of the organic acid include citric
20 acid, stearic acid, succinic acid and tocopherol succinate. Examples of the anionic surfactant include dodecyl sulfate. Other preparation conditions are similar to those of the ion-exchange method.

"Hybrid materials" as the term is used herein do not refer to mere mixtures, but rather refer to such materials made based on chemical bonding forces between
25 constituents. For example, in the case of cationic layered metal hydroxides and anionic raw materials for cosmetic active components, electrostatic interactions between the components having the counter ions act as the main chemical bonding force. Examples of a method based on such chemical bonding force include the coprecipitation method and the ion-exchange method. In the adsorption method,
30 van der Waals force between the first introduced material, e.g. tocopherol succinate, and the later introduced material, e.g. retinol, is the main chemical bonding force. However, these are given by way of explanation and are not intended to limit the scope of the invention. Thus, in real conditions, electrostatic interactions and van der Waals force may coexist partly depending on constituents or preparation
35 conditions.

Because the surface of the hybrid materials obtained as above is composed

of inorganic materials, coagulation occurs during the preparation process through coprecipitation in a solution, or during the drying step, which causes deterioration of particle uniformity, mixability, and dispersibility and dispersion stability.

One of the solutions to such problems includes a method of coating the surface of the obtained hybrid materials. Surface coating materials may be inorganic, organic or polymer materials which are harmless to the human body.

Inorganic coating materials include silica (SiO_2), titania (TiO_2) and alumina (Al_2O_3), but silica is most preferable in view of surface property modification performance, preparation process, cost of raw materials, etc. Raw materials of the inorganic coating materials include metal salts, e.g. SiCl_4 , TiCl_4 , $\text{Si}(\text{OC}_2\text{H}_5)_4$, $\text{Si}(\text{OCH}_3)_4$, $\text{Ti}(\text{OC}_3\text{H}_7)_4$ and $\text{Al}(\text{OC}_3\text{H}_7)_3$. Preferably, alkoxides such as $\text{Si}(\text{OC}_2\text{H}_5)_4$, $\text{Si}(\text{OCH}_3)_4$, $\text{Ti}(\text{OC}_3\text{H}_7)_4$ and $\text{Al}(\text{OC}_3\text{H}_7)_3$ is used. Surface coating using the inorganic coating materials are characterized by hydrolysis reaction between the hybrid materials and the coating raw materials. If necessary, a basic or acidic catalyst may be added.

In addition to the inorganic coating materials, various neutral surfactants having hydrophilic or lipophilic properties, and polymers such as polyvinyl alcohol, polyethylene glycol, cellulose and polymethyl methacrylate can be used alone or in combination. Alternatively, silane compounds and polymers can be used simultaneously for coating.

The present invention also provides cosmetics comprising the aforementioned hybrid materials of raw materials for cosmetics, active components and layered metal hydroxides. The cosmetics may comprise mediums and other additives conventionally used in cosmetics.

The hybrid materials according to the present invention show several properties distinguished from the materials prepared by prior stabilization methods. First, the hybrid materials according to the present invention have the particle size of tens of nm to several μm . If necessary, a drying step may be omitted to give products in the form of a suspension, colloid or slurry. Furthermore, products in the form of granules can be easily prepared to satisfy the conditions required for various formulation/addition processes. Second, since the hybrid material capsule consists of inorganic materials, stability is superior to prior capsules for stabilization. Furthermore, effects of blocking heat, light, moisture, oxygen, etc. are also excellent. In particular, degeneration of active components by heat or light is remarkably reduced. Third, the amount of active components per unit mass of the hybrid material capsule is very high compared to organic capsules. Fourth, the hybrid

materials according to the present invention is biocompatible inorganics which show less adverse effects such as toxicity and irritation compared to prior materials for stabilization. Fifth, mixability, dispersibility, dispersion stability, touch, convenience, etc. of the hybrid materials can be improved by surface coating. Sixth, cosmetics containing the hybrid materials as raw materials show additional functions such as blocking ultraviolet rays by the inorganic capsule materials and removing skin wastes by adsorption properties of the hybrid materials. Seventh, the preparation process is relatively simple and economical.

Furthermore, cosmetics containing the hybrid materials according to the present invention show less deterioration properties such as skin whitening, removal of keratin, preventing wrinkles, removing spots, skin moisturizing, blocking ultraviolet rays and removing skin wastes that are inherent to each active component due to the aforementioned properties of the hybrid materials.

The present invention will be described below in further detail by Examples. However, the Examples are only illustrative, and not intended to limit the scope of the invention.

EXAMPLES

Example 1

450 g of $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 6\text{H}_2\text{O}$ was completely dissolved in 4 liters of distilled water and then 55 g of sodium ascorbate was added and again completely dissolved. 500 ml of 2 N NaOH solution was slowly added to the mixed solution to make pH of the solution 6.5, and precipitation was induced with stirring for 12 hours. The reaction temperature was kept below 10°C . Upon completion of the reaction, the precipitate and the filtrate were separated by centrifuge, and the precipitate was washed with distilled water 5 times or more to remove impurities. Then, the precipitate in slurry was dried using a freezing dryer to obtain the powder in fine particles. Figure 1 is an X-ray diffraction pattern of the obtained hybrid material of zinc hydroxide and vitamin C. Analysis of the diffraction pattern shows that the hybrid material is a layered material in the form of hydrozincite substituted by vitamin C. Figure 2 is a field emission scanning electron microscopy (FE-SEM) micrograph of the hybrid material. The scan shows that the hybrid material consists of spherical fine particles of 100 – 200 nm mono-dispersion.

Example 2

The same conditions as those of Example 1 were used to produce a hybrid material of zinc hydroxide and vitamin C as obtained in Example 1, except using ascorbic acid instead of sodium ascorbate as the starting material.

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Example 3

The same conditions as those of Example 1 was used to produce a hybrid material, except preparing a metal aqueous solution by dissolving 620 g of $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ instead of Zn-acetate in 4 liters of distilled water and using 175 g of ascorbic acid as the starting material of vitamin C. The hybrid material of zinc hydroxide and vitamin C prepared in the present Example showed an X-ray diffraction pattern similar to that of Example 1. This means that the product is also a layered compound having the crystal structure of hydrozincyte. Figure 3 is a FE-SEM diagram of the vitamin hybrid material produced from $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ as the starting material, which shows the size and the structure of particles of the hybrid material. This shows that the average particle size of the hybrid material is about 50 ~ 100 nm.

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Example 4

A hybrid material of L-ascorbic acid (vitamin C) and metal hydroxide in the form of hydrozincyte was synthesized using the ion-exchange method. 620 g of $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ was completely dissolved in 4 liters of distilled water. Then, 500 ml of 2 N NaOH solution was slowly added to the solution to make pH of the solution 4.5. Precipitation was induced with stirring for 12 hours. Reaction was performed at room temperature with continuous flow of nitrogen gas into the reaction solution. After 12 hours of reaction, the precipitate and the filtrate were separated using a centrifuge. The precipitate was washed with distilled water 5 or more times to remove impurities. After washing, the precipitate in slurry was freeze-dried to obtain powder in fine particles. X-ray diffraction showed that the powder in fine particles obtained had the crystal structure of $\text{Zn}_5(\text{OH})_8(\text{NO}_3)_2 \cdot n\text{H}_2\text{O}$ in the form of hydrozincyte. 0.1 g of thus obtained inorganic capsule material and 0.3 g of ascorbate were added to 100 ml of distilled water. Then, the ion-exchange reaction was performed with stirring to introduce ascorbate ions into the inorganic capsule material. The reaction solution was kept at about pH 7, and the reaction was performed for 24 hours with stirring. After completion of the ion-exchange reaction, a solid product was separated through a centrifuge, washed several times with distilled water, and freeze-dried to obtain a hybrid material of zinc hydroxide

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and vitamin C. The resulting hybrid material was found to have the same crystal structure as that of the hybrid material prepared in Example 3, using X-ray diffraction.

Example 5

5 6 g of $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 6\text{H}_2\text{O}$ and 0.4 g of NaOH were completely dissolved in 0.3 liter of 25 % ethanol solution. Then, 0.75 g of a retinic acid was added to the solution and completely dissolved. Precipitation was induced with stirring for 12 hours. Next, a hybrid material of zinc hydroxide and retinic acid (vitamin A) in fine powder with the other conditions kept the same as those of Example 1. Figure 4 is
10 an X-ray diffraction pattern of thus obtained hybrid material of zinc hydroxide and vitamin A. Analysis of the diffraction pattern shows that the hybrid material is a layered material substituted by vitamin A in the form of hydrozincyte. From element analysis and heat analysis of the metal ion and the organic, the molecular weight of the hybrid material was $\text{Zn}_5(\text{OH})_8(\text{vitamin A})_2 \cdot n\text{H}_2\text{O}$. Figure 5 is a field
15 emission scanning electron microgram, which shows that the hybrid material is composed of mono-disperse spherical fine particles of about 50 nm.

Example 6

A hybrid material of Zn/Al-LDH and D- α -tocopherol succinate-retinol
20 (vitamin E derivative) was synthesized using the adsorption method. 4.09 g of ZnCl_2 and 1.33 g of AlCl_3 were dissolved in 400 mL of decarboxylated distilled water, and then 70 mL of an aqueous solution of D- α -tocopherol succinate (2.19 g) was added to the mixture. A coprecipitation reaction was performed by adding 1 N aqueous solution of NaOH to the mixed solution with stirring under the atmosphere
25 of nitrogen to the final pH 7.5. The reaction was continued at room temperature for 12 hours. Then, the precipitate was separated from the reaction mixture by a centrifuge. The precipitate was washed 5 times and freeze-dried to obtain a hybrid material of vitamin E and metal hydroxide. An X-ray diffraction pattern of thus
30 obtained hybrid material showed that the distance between the layers was 47.0 Å, and that the hybrid material is vitamin E derivative arranged in the form of double-layers between the hydroxide layers. 1 g of thus obtained hybrid material was dispersed enough in 50 mL of ethanol. Then, 1 g of vitamin A (retinol) was dissolved and reacted under nitrogen atmosphere with stirring for 24 hours. After
the reaction, solid/liquid was separated using a centrifuge, washed several times
35 using ethanol, and dried to obtain a hybrid material of [vitamin E]-[vitamin A]-[inorganic]. Figure 6 is an X-ray diffraction pattern of thus obtained hybrid

material, which shows the distance between the layers is 72.1 Å, and the diffraction pattern moved to lower angle compared to that before adsorbing retinol. This is a phenomenon which occurs when vitamin A molecules are inserted between vitamin E derivatives that have already been inserted between the layers.

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Example 7

A hybrid material of lactic acid (α -hydroxy acid) and metal hydroxide in the form of hydrozincyte was synthesized using the coprecipitation method. 23 g of $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ was completely dissolved in 300 mL of distilled water, and then 9 g of lactic acid was added and completely dissolved. 100 ml of 0.5 N NaOH solution was slowly added to the mixed solution to make pH of the solution 6. Then, precipitation was induced with stirring under nitrogen atmosphere for 12 hours. After completion of the reaction, the precipitate and the filtrate were separated using a centrifuge, and the precipitate was washed 5 times or more with distilled water to remove impurities. Then, the precipitate in slurry was dried under vacuum to obtain powder in fine particles. Analysis of an X-ray diffraction pattern shows that the hybrid material is a layered hydroxide containing anions of the lactic acid between the layers. Figure 7 is a field emission scanning electron microgram of the hybrid material, which shows that the hybrid material is composed of fine dispersed mono particles of 100 to 200 nm.

Example 8

A hybrid material of Zn/Al-LDH and citric acid was synthesized using the coprecipitation method. 465 g of $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and 193 g of $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ were completely dissolved in 3 L of distilled water, and then 74 g of citric acid was added and completely dissolved. 4 N NaOH solution was slowly added to the mixed solution to make pH of the solution 8. Then, precipitation was induced with stirring for 12 hours. After completion of the reaction, the precipitate was separated, washed and dried in the same manner as in Example 7 to obtain powder of the hybrid material in fine particles. Figure 8 is an X-ray diffraction pattern of the obtained hybrid material of Zn/Al-LDH and citric acid. Analysis of the pattern shows that the hybrid material is a layered double hydroxide in the form of hydrotalcite containing anions of the citric acid between the layers.

Example 9

A hybrid material of kojic acid and hydrozincyte-form metal hydroxide was

synthesized using the coprecipitation method. 10.34 g of kojic acid ($C_6H_6O_4$; molecular weight = 142.11 g/mol) was completely dissolved in a mixed solution of 100 mL of decarboxylated water and 100 mL of ethanol, and then 0.6 N NaOH aqueous solution was added to make pH of the solution 9~10. 18 g of $Zn(NO_3)_2 \cdot 6H_2O$ was dissolved in 100 mL of decarboxylated water in another vessel, and then this aqueous metal solution was added to the solution containing the kojic acid. At the same time, pH of the reaction solution was adjusted to about 9 to 10 using 0.6 N NaOH aqueous solution. After completion of the titration, the solution was reacted at room temperature for 20 hours. After completion of the reaction, the precipitate was separated from the filtrate using a centrifuge, and washed using 5 times or more using distilled water to remove impurities. Then, the precipitate in slurry was dried under vacuum to obtain powder in fine particles. Figure 9 is an X-ray diffraction pattern of thus obtained hybrid material, which shows that a distance between the layers is 14.6 Å. This means that the hybrid material is a layered hydroxide containing anions of the kojic acid between the layers.

Example 10

The same conditions as those of Example 7 were used to produce a hybrid material, $Zn_5(OH)_8(\text{salicylic acid})_2 \cdot 4H_2O$, except that 2.3 g of salicylic acid ($C_7H_6O_3$) was used, and the solution was kept at pH 6. An X-ray diffraction pattern of the hybrid material is shown in Figure 10.

Example 11

37 g of $Zn(NO_3)_2 \cdot 6H_2O$ was completely dissolved in 120 ml of decarboxylated distilled water, and then 10 g of indol-3-acetic acid was completely dissolved in 15 ml of decarboxylated distilled water and 120 ml of ethanol. Then, 50 ml of 1.6 N NaOH solution was slowly added to the mixed solution to make pH of the solution 6.7, and precipitation was induced by stirring for 6 hours. The reaction temperature was kept below 10 °C. The precipitate was sufficiently washed and dried to obtain powder in fine particles. An X-ray diffraction pattern of thus obtained hybrid material is shown in Figure 11, and a scanning electron microgram is shown in Figure 12. The X-ray pattern shows that a distance (d_{001}) between the layers is 18.2 Å, and this means that molecules of indol-3-acetic acid are stabilized between the layers of the hydroxide, and that the produced hybrid material is fine particles of 100 to 200 nm.

Example 12

100 g of the precipitate prepared in Example 3 in slurry was added to a mixed solution of 100 g of tetraethoxysilicate (TEOS; $\text{Si}(\text{OC}_2\text{H}_5)_4$) dispersed in 400 ml of ethanol, and stirred for 1 hour. Then, 100 ml of distilled water was added to the reaction solution to perform a coating reaction. The coating reaction was carried out at room temperature for 24 hours. After completion of the reaction, the solid product was separated using a centrifuge, and then washed two times using ethanol. Thus obtained solid material was dried in a vacuum drier for 24 hours to give powder in fine particles. Figure 13 is a field emission scanning electron microgram, which shows that silica particles of 20 to 30 nm is coated onto the surface of the hybrid material, and many pores have been formed between the particles.

Example 13

A hybrid material of hydrozincite-type metal hydroxide and vitamin C was synthesized, and without separate separation or washing step, tetraethoxysilicate (TEOS) was directly added to coat the surface of the hybrid material with silica (SiO_2). 42.70 g of ascorbic acid was completely dissolved in 60 mL of decarboxylated water, and then an aqueous solution of 185.92 g of $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ dissolved in 300 mL of decarboxylated water was added in drops. At the same time, 240 mL of 1.6 N NaOH aqueous solution was titrated, and reacted for 5 hours to obtain the hybrid material of vitamin C and hydrozincite in coprecipitate. After completion of the reaction, without washing step, a mixed solution of 394.68 g of TEOS and 2.4 L of ethanol was slowly added in drops to the suspension of the above hybrid material. Then, the coating reaction was carried out for a further 12 hours. After completion of the reaction, the reaction mixture was centrifuged, washed with ethanol 4 times, and dried to obtain the hybrid material coated with silica. Figure 14 is a scanning electron microgram of thus obtained particles, which shows that particles having a size of about 0.5 μm have been formed.

Example 14

The hybrid material obtained in Example 3 in slurry was freeze-dried to produce powder. The powder was surface-coated with silica in the same way as in Example 12. In doing this, 5 g of the dried hybrid material was added to a mixed solution of 20 g of TEOS dispersed in 80 ml of ethanol, and stirred for 1 hour. Then, 20 ml of distilled water was added to the reaction solution, and the coating

reaction was carried out at room temperature for 24 hours. After completion of the coating reaction, the product was washed and dried in the same way as in Example 12 to obtain a product in particles. Figure 15 is an electron microgram of thus coated material. This shows that, like Example 14, silica particles of tens of nm
5 have been coated onto the surface of the hybrid material and many pores have been formed between the particles.

Example 15

100 g of the slurry-phase precipitate obtained in Example 3 was added to a
10 mixed solution of 100 g of tetramethoxysilicate (TMOS; $\text{Si}(\text{OCH}_3)_4$) dispersed in 400 ml of methanol. After stirring for 1 hour, 100 ml of distilled water was added and the coating reaction was carried out at room temperature for 24 hours. After completion of the reaction, a solid product was separated using a centrifuge, and washed two times using ethanol. Then, the obtained solid material was dried in a
15 vacuum drier for 24 hours to obtain powder in fine particles. Figure 16 is a field emission scanning electron microgram of thus coated hybrid material, which shows that silica particles of 20 to 30 nm have been coated onto the surface of the hybrid material and many pores have been formed between the particles when coated using TMOS as well.

Example 16

100 g of the slurry-phase precipitate obtained in Example 3 was added to a mixed solution of 100 g of tetraethoxysilicate (TEOS; $\text{Si}(\text{OC}_2\text{H}_5)_4$) dispersed in 400 ml of ethanol, and then the coating reaction was carried out with stirring. At this
25 time, a small amount of 1 N HCl solution was added in drops to adjust pH of the reaction solution to 5, thereby inducing the coating reaction under a weakly acidic condition. The coating reaction was carried out at room temperature for 24 hours. After completion of the reaction, a solid product was separated using a centrifuge, and washed two times using ethanol. Then, the obtained solid material was dried in
30 a vacuum drier for 24 hours to obtain powder in fine particles. Figure 17 is a field emission scanning electron microgram of thus coated hybrid material, which shows that uniformity of the surface coating has been improved using an acid catalyst during the coating reaction.

Example 17

A hybrid material of L-ascorbic acid (vitamin C) and Zn/Al-LDH was

synthesized using the ion-exchange method. Using the metal salts $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ as starting materials, an aqueous solution containing the metal ions Zn(II) and Al(III) in suitable mole ratio ($\text{Zn}/\text{Al} = 2$) was prepared. Then, 0.2 N NaOH solution was added in drops to the solution to obtain a layered double hydroxide with nitric acid anions (NO_3^-) inserted between the layers. The precipitate was washed enough and dried to obtain a LDH precursor in the form of fine powder. This was mixed with an aqueous solution containing vitamin C to allow ion-exchange with nitric acid anions between the layers, thereby obtaining a hybrid material of vitamin C and the layered metal hydroxide. In doing this, a mole ratio of vitamin C to the inorganics, $[\text{vitamin C}]/[\text{Al}]$, was 1.5, and the reaction was carried out with stirring at room temperature for 24 hours. An X-ray diffraction analysis of thus obtained hybrid material of vitamin C and Zn/Al-LDH shows that the distance between the layers is 10.5 \AA , and that vitamin C molecules have been stabilized as a single layer parallel with the layer of the metal hydroxide. The hybrid material thus obtained by the ion-exchange was surface-coated with silica using TEOS as a precursor. The coating reaction was carried out by adding the hybrid material precipitate in slurry to a mixed solution of 100 g of TEOS dispersed in 400 ml of ethanol and stirring for 1 hour, followed by adding 100 ml of distilled water to the reaction solution. The subsequent procedures were the same as those in Example 12. Figure 18 is a scanning electron microgram, which shows that the hybrid material in the form of LDH as well has been surface-coated with silica particles.

Example 18

The slurry-phase precipitate obtained in Example 3 was washed two times using ethanol, and solid/liquid separated using a centrifuge to obtain a gel. 40 g of thus obtained gel was dispersed in 80 mL of ethanol, and stirred under nitrogen atmosphere. Also, 10.7 g of TMOS and 2.9 g of γ -methacryloxypropyl-trimethoxysilane were added to 93 mL of ethanol. To this solution, 1.5 mL of 0.1 N HCl aqueous solution was added and hydrolyzed at ambient temperature for 30 minutes. Then, this solution was added to the above gel dispersion and stirred at ambient temperature for 30 minutes for reaction. 2.25 g of 2,2'-azobisisobutyronitrile was added to this solution and after 5 minutes 29.2 mL of methylmethacrylate was added. After reaction at 60°C with stirring for 4 hours, solid was separated using a centrifuge, washed two times with ethanol, and dried under vacuum to obtain a hybrid material coated with silica-PMMA. A scanning

electron microgram of thus obtained hybrid material is shown in Figure 19. This shows that very uniform and mono-disperse hybrid material has been obtained.

Example 19

5 5 g of the particulate material obtained in Example 12 was introduced into a 250 mL conical flask, and sealed with a silicon stopper under nitrogen-filled condition. Then, 15 g of absolute methanol and 2.5 g of methyl methacrylate were added using a syringe, and stirred. Also, 0.1 g of 2,2'-azobisisobutyronitrile (AIBN) was dissolved in 3.75 g of absolute methanol, and this solution was added to the
10 above reactor by a syringe. To absorb the internal pressure, a rubber balloon was put up in the reactor. The solution was reacted with stirring at 60 °C for 24 hours. After completion of the reaction, solid was separated using a centrifuge, washed two times with ethanol, and dried under vacuum to obtain PMMA-coated powder. A scanning electron microgram of the PMMA-coated hybrid material is shown in
15 Figure 20.

Example 20

 The particulate material obtained in Example 12 was surface-coated using cellulose. 10 g of the particulate material and 30 g of Tween 20, a neutral surfactant,
20 was added into 200 mL of distilled water and dispersed well. 5 g of ethyl cellulose dissolved in methylene chloride was added to the solution containing the silica-coated vitamin hybrid material and emulsified. This emulsion was stirred using a homogenizer until methylene chloride is evaporized up, and centrifuged. The obtained gel was washed three times with distilled water, and then dried under
25 vacuum to obtain powder coated with ethyl cellulose. A scanning electron microscopy of thus obtained hybrid material is shown in Figure 21. This shows that coated particles having a good mono-dispersity have been obtained by the surface treatment using ethyl cellulose.

Example 21

 The particulate material obtained in Example 12 was surface-coated using polymethylmethacrylate (PMMA). 2.5 g of the particulate material was added into 200 mL of aqueous solution containing 1.0 w/v% of polyvinyl alcohol, and dispersed well. Then, 5 g of polymethylmethacrylate dissolved in 300 mL of methylene
35 chloride was added to the dispersion of the vitamin hybrid material, and emulsified. This emulsion was stirred using a homogenizer until methylene chloride is vaporized,

then, centrifuged. The obtained gel was washed three times with distilled water, and then dried under vacuum to obtain powder coated with polymethylmethacrylate. A scanning electron microscopy of thus obtained hybrid material is shown in Figure 22. This shows that uniform spherical coating particles have been obtained.

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Example 22

$\text{Zn}_5(\text{OH})_8(\text{C}_{10}\text{H}_8\text{NO}_2)_2 \cdot n\text{H}_2\text{O}$ obtained in Example 11 was surface-coated using silica. The coating reaction was carried out for 12 hours by stirring 1 g of the above hybrid material in 10 ml of an aqueous solution for 1 hour to dispersion, and then adding in drops 1 g of TEOS ($\text{Si}(\text{OC}_2\text{H}_5)_4$) dissolved in 10 ml of ethanol. Figure 23 is a scanning electron microgram of the hybrid material coated with silica, which shows that silica of 20 to 30 nm has been uniformly coated onto the surface of the hybrid material.

15 **Example 23**

Storage stability in an aqueous solution was evaluated for the hydrozincite-type vitamin C hybrid material coated with silica, obtained in Example 12. From the analysis of the vitamin C hybrid material powder obtained in Example 12, the content of vitamin C was 25 wt%. 400 mg of the powder was added to 10 mL of decarboxylated distilled water, and sealed. Then, while storing the hybrid material in a constant-temperature oven at 42 °C, the content of vitamin C was periodically examined using HPLC. In order to compare stability, 100 mg of pure vitamin C (L-ascorbic acid) was stored in the same condition, and a change of the vitamin content was examined. Figure 24 is a graph showing the changes of the vitamin C contents depending on the storage period. This shows that the vitamin C content of the hybrid material according to the present invention did not show significant change, while vitamin C was destroyed very rapidly in the pure vitamin C.

25 **Example 24**

Sustained-release property was evaluated for the vitamin C hybrid material obtained in Examples 3 and 12. For this, 0.00738 mol of the hybrid material powder was added to 25 ml of 0.8 % NaCl aqueous solution, and with stirring uniformly at 37 °C, the released amount of vitamin C depending on time was measured through an absorption intensity at 265 nm of ultraviolet absorption spectrum. The results are shown in Figure 25. As seen from Figure 25, in the present experimental condition, vitamin C was released continuously as time lapsed,

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which shows that the vitamin hybrid material according to the present invention has a good sustained-release property.

Example 25

- 5 A sustained-release property of indol-3-acetic acid molecules was evaluated for the indol-3-acetic acid hybrid material obtained in Example 11. For this, 0.5 g of the hybrid material was added to 25 ml of 0.8 % NaCl aqueous solution, and with stirring uniformly at 37 °C, the released amount of vitamin C depending on time was measured through high performance liquid chromatography (HPLC). Figure 26 is a
- 10 graph showing a release rate of indol-3-acetic acid measured in this way, which shows that indol-3-acetic acid anions stabilized between the layers were slowly released by the ion-exchange with Cl⁻ ions in the saline solution.

CLAIMS

1. A hybrid material comprised of an active component for raw materials for cosmetics and a layered metal hydroxide, the active component selected from the group consisting of vitamin, lactic acid, citric acid, kojic acid, indol-3-acetic acid and salicylic acid.
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2. The hybrid material according to Claim 1, wherein the active component for raw materials for cosmetics is a pure vitamin of vitamin A, vitamin C or vitamin E, or a derivative thereof.
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3. The hybrid material according to Claim 1, wherein the layered metal hydroxide is a hydroxy double salt.
4. The hybrid material according to Claim 1, wherein the layered metal hydroxide is a layered double hydroxide.
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5. The hybrid material according to Claim 1, wherein the surface has been coated.
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6. The hybrid material according to Claim 5, wherein the surface has been coated with silica.
7. The hybrid material according to Claim 5, wherein the surface has been coated with at least one coating material selected from the group consisting of polyvinyl alcohol, polyethylene glycol, cellulose, polymethyl methacrylate and silane.
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8. A method for preparing the hybrid material according to any one of Claims 1 to 7, comprising:
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dissolving an active component for raw materials for cosmetics and starting materials for a layered metal hydroxide in an aqueous solvent or a mixed solvent, thereby causing precipitation.
9. The method according to Claim 8, wherein the active component for raw materials for cosmetics is selected from the group consisting of retinic acid, L-
- 35

ascorbic acid, tocopherol succinate, lactic acid, citric acid, salicylic acid, kojic acid and indol-3-acetic acid.

10. A method for preparing the hybrid material according to any one of Claims 1
5 to 7, comprising:

(1) reacting starting materials for a layered metal hydroxide in a basic solution to form a layered metal hydroxide salt; and

(2) dispersing the said layered metal hydroxide salt in an aqueous solvent or a mixed solvent, and dissolving an active component for raw materials for cosmetics,
10 thereby causing ion-exchange reaction.

11. A method for preparing the hybrid material according to any one of Claims 1 to 7, comprising:

(1) reacting starting materials for a layered metal hydroxide and an organic
15 acid or an anionic surfactant in an aqueous solvent or a mixed solvent to form a coprecipitate of an organic-inorganic layered hybrid material; and

(2) dispersing the organic-inorganic layered hybrid material in an aqueous solvent or a mixed solvent, and adding an active component for raw materials for cosmetics, thereby adsorbing the active component onto the organic-inorganic
20 layered hybrid material.

12. The method according to Claim 11, wherein the active component is selected from the group consisting of retinol, retinal, retinyl acetate, retinyl palmitate, α -tocopherol and tocopheryl acetate.

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13. The method according to Claim 11, wherein the starting materials for a layered metal hydroxide are reacted with the organic acid selected from the group consisting of citric acid, stearic acid, succinic acid and tocopherol succinate, to form a coprecipitate.

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14. The method according to Claim 11, wherein the starting materials for a layered metal hydroxide are reacted with an anionic surfactant such as dedecyl sulfate, to form a coprecipitate.

35 15. The method according to any one of Claims 8 to 14, further comprising coating the surface of the hybrid material.

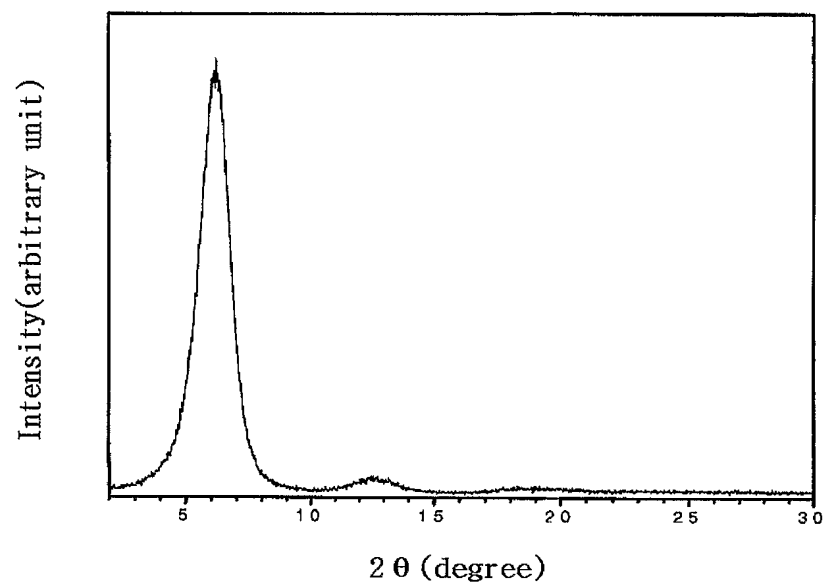
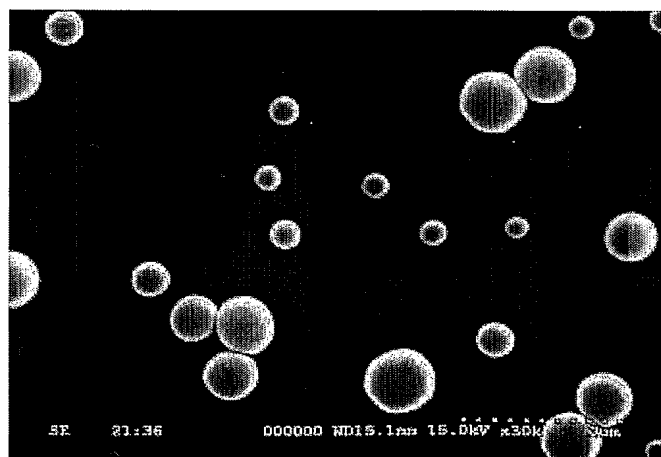
16. The method according to Claim 15, wherein the material for coating is silica.

17. The method according to Claim 15, wherein the material for coating is at
5 least one selected from the group consisting of polyvinyl alcohol, polyethylene glycol, cellulose, polymethyl methacrylate and silane.

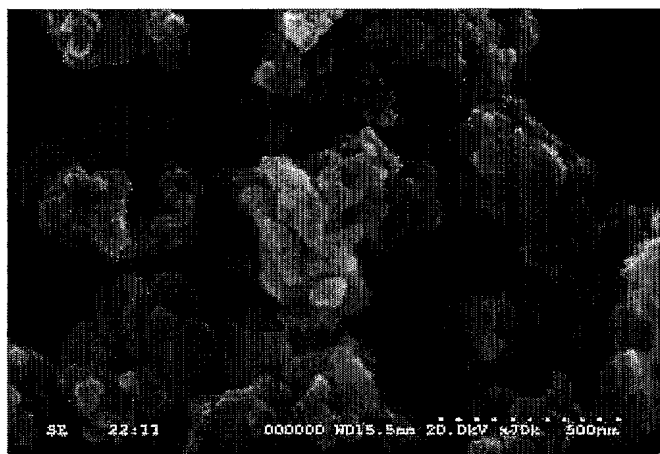
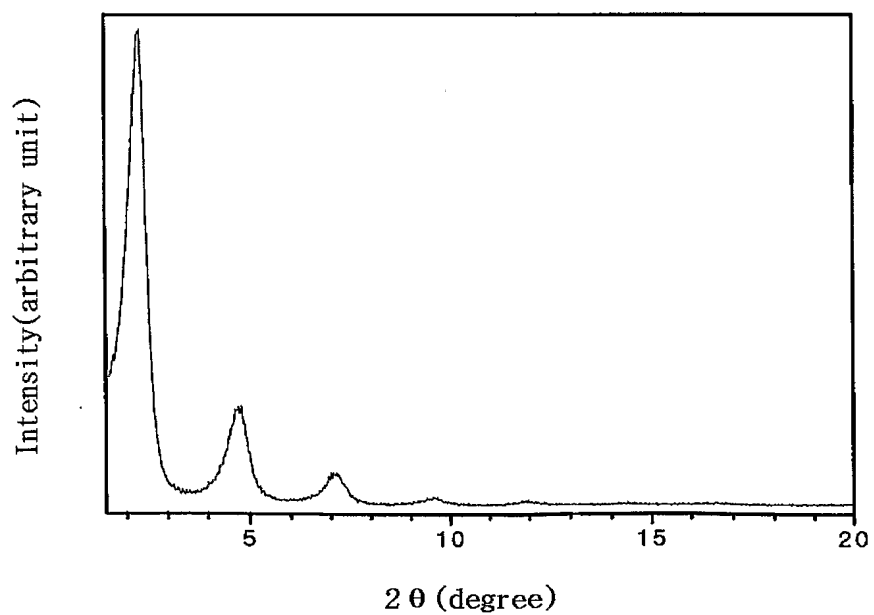
18. Cosmetics comprising the hybrid material according to any one of Claims 1
to 7.

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Fig. 1**Fig. 2**

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Fig. 3**Fig. 4**

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Fig. 5

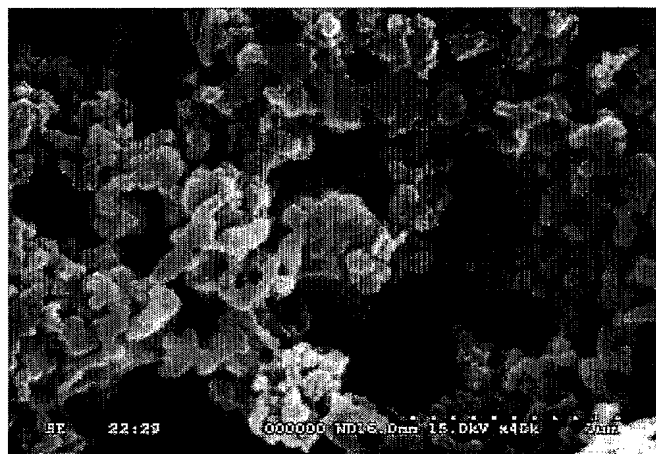
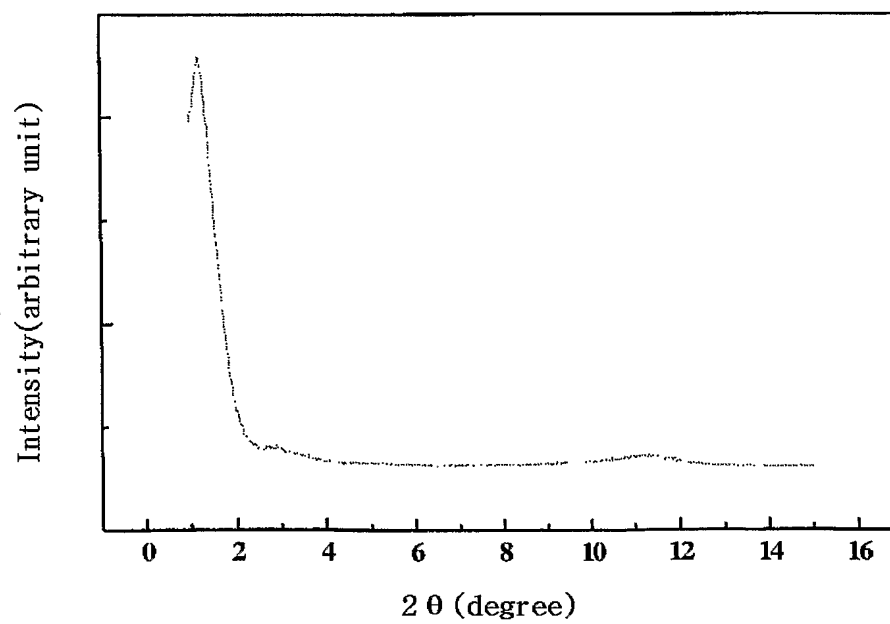


Fig. 6



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Fig. 7

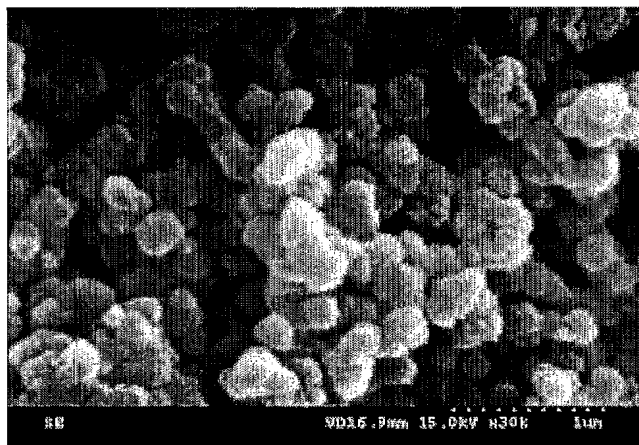
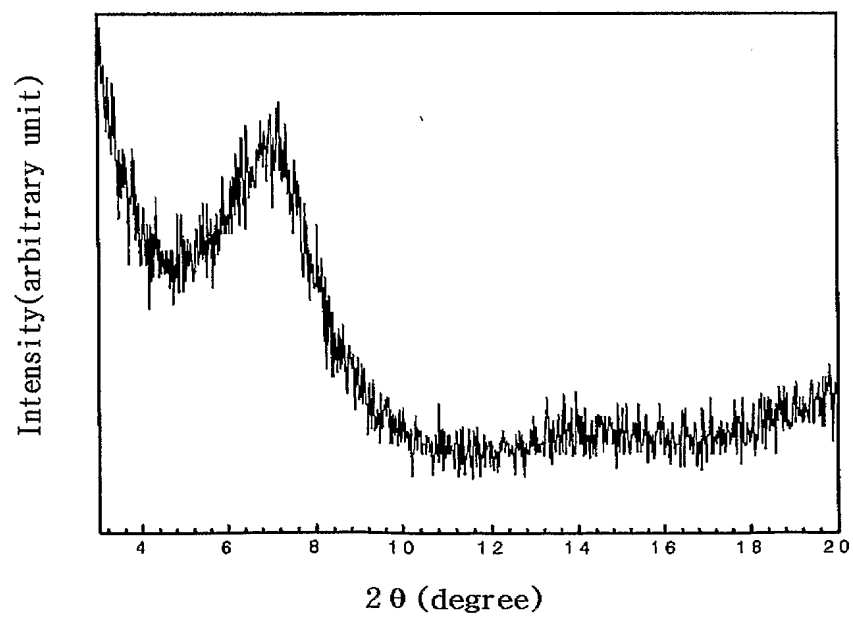
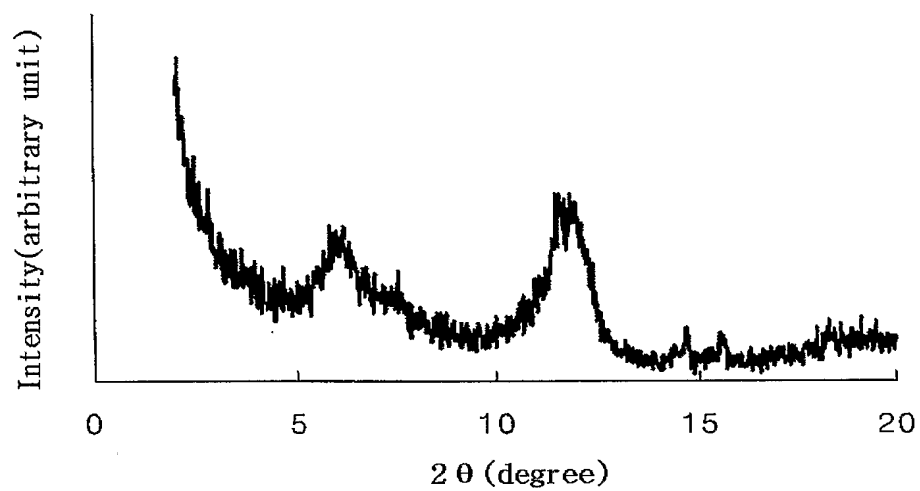
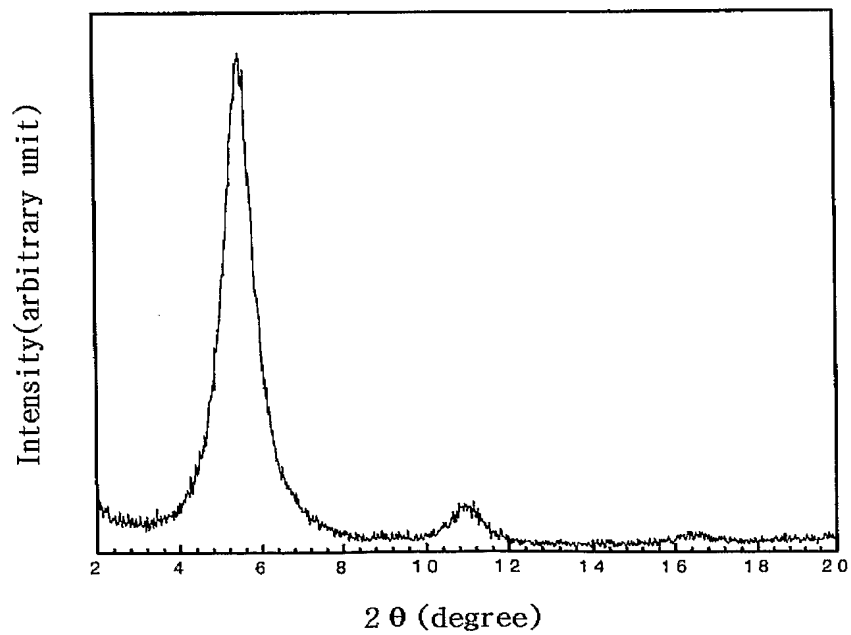


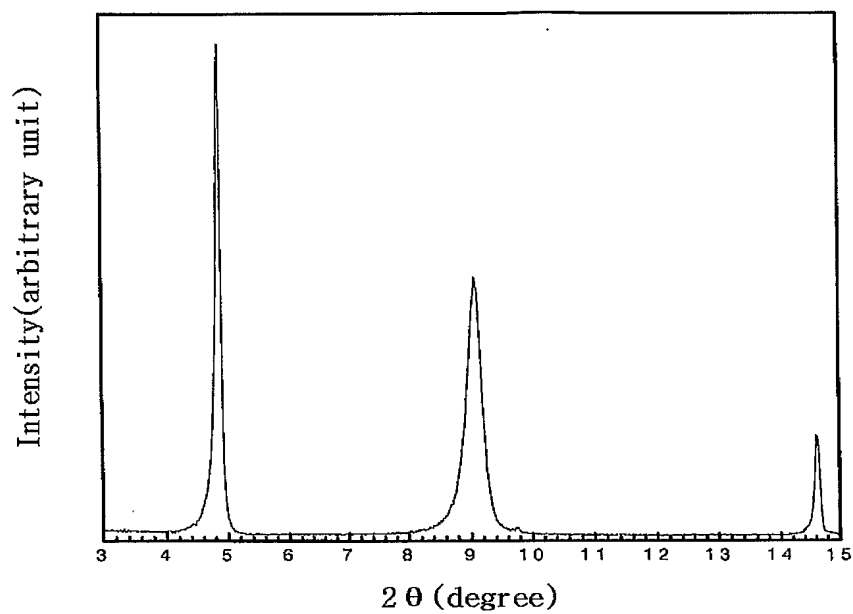
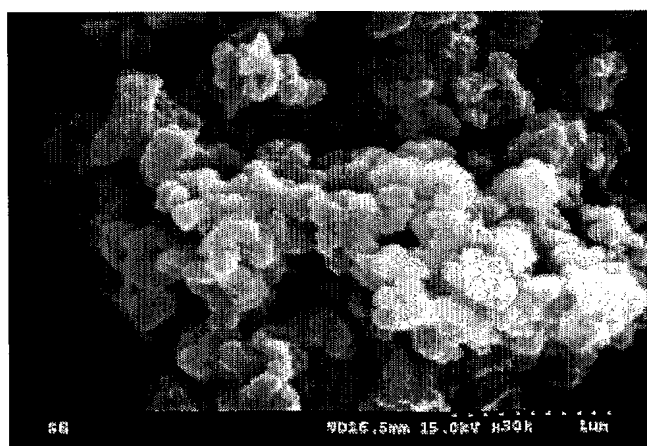
Fig. 8



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Fig. 9**Fig. 10**

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Fig. 11**Fig. 12**

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Fig. 13

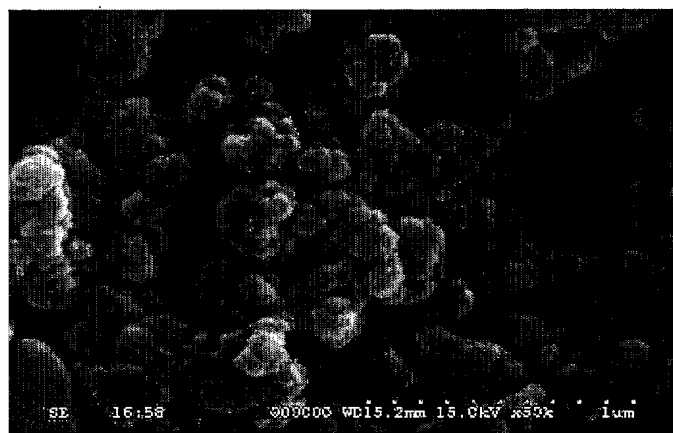
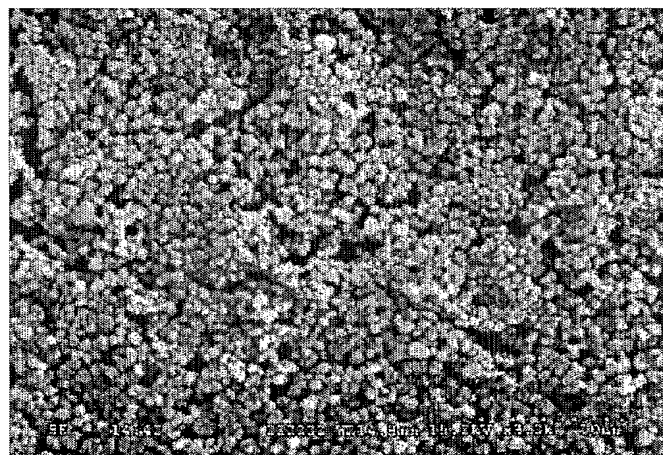


Fig. 14



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Fig. 15

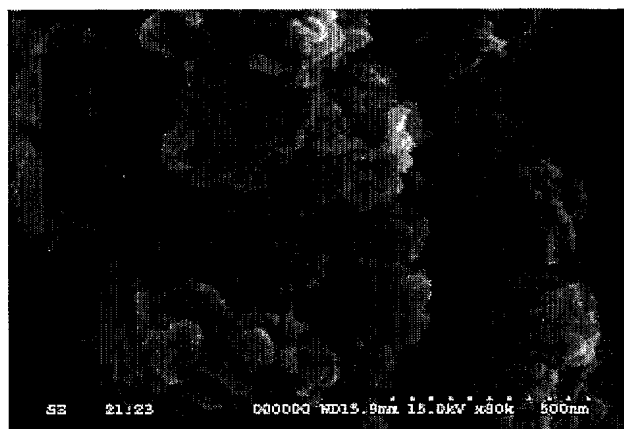
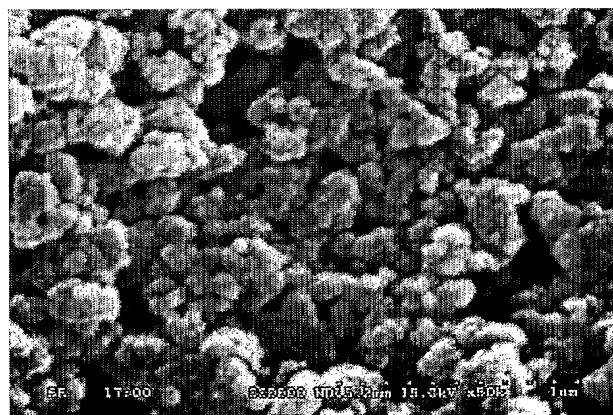


Fig. 16



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Fig. 17

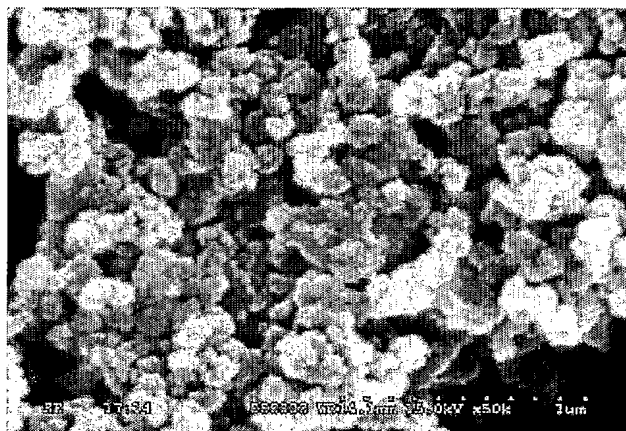
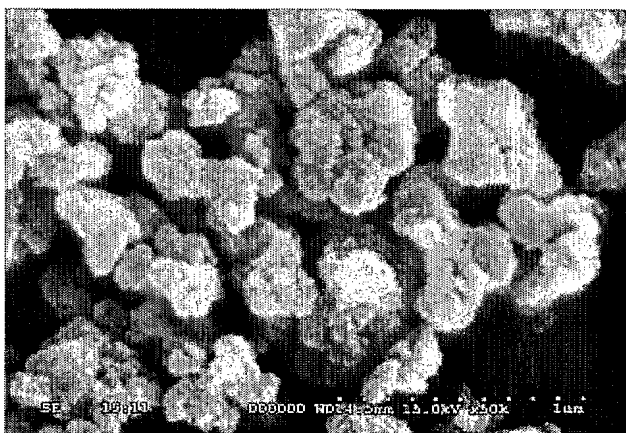


Fig. 18



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Fig. 19

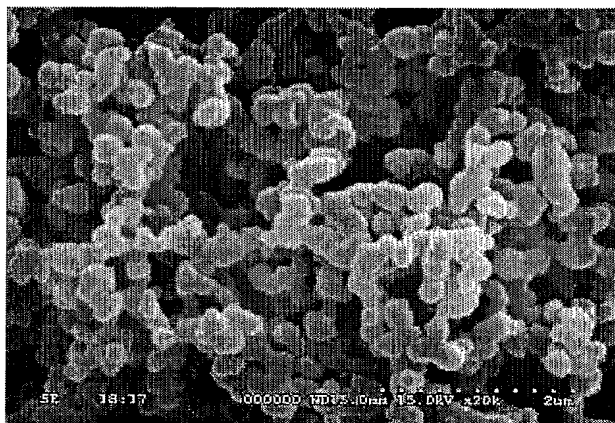
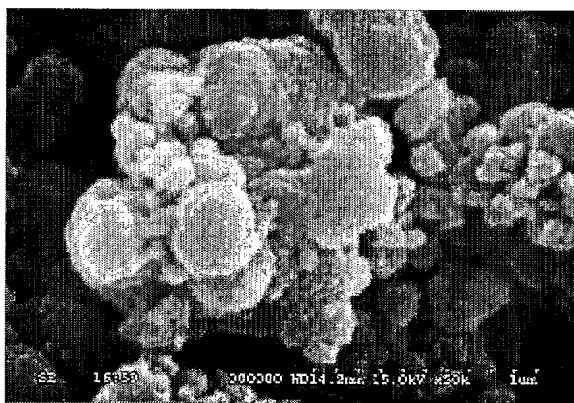


Fig. 20



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Fig. 21

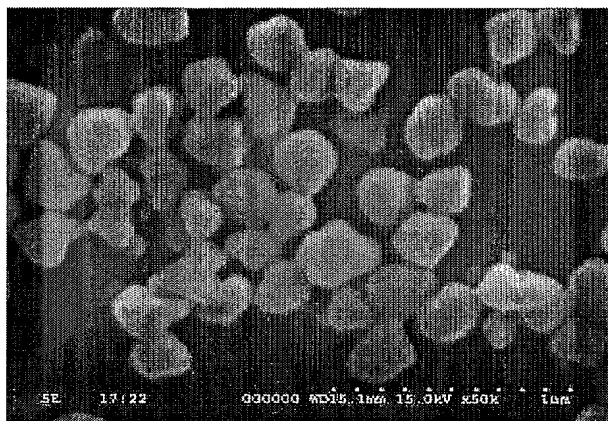
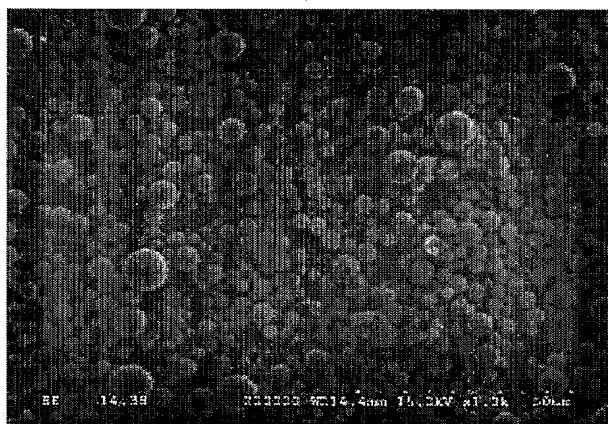


Fig. 22



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Fig. 23

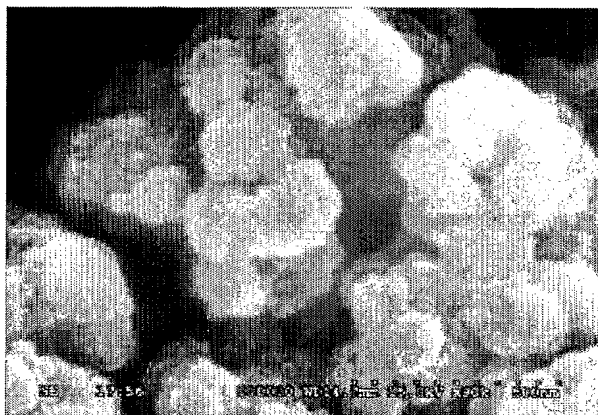
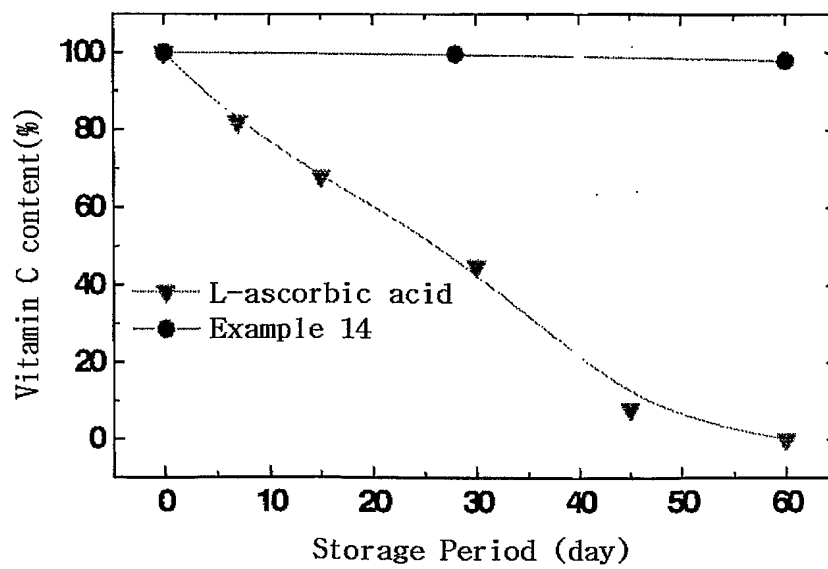
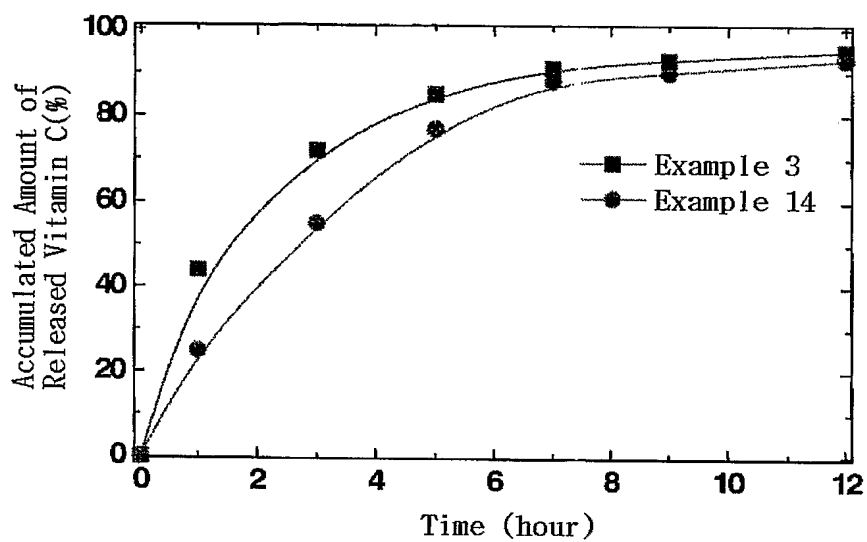
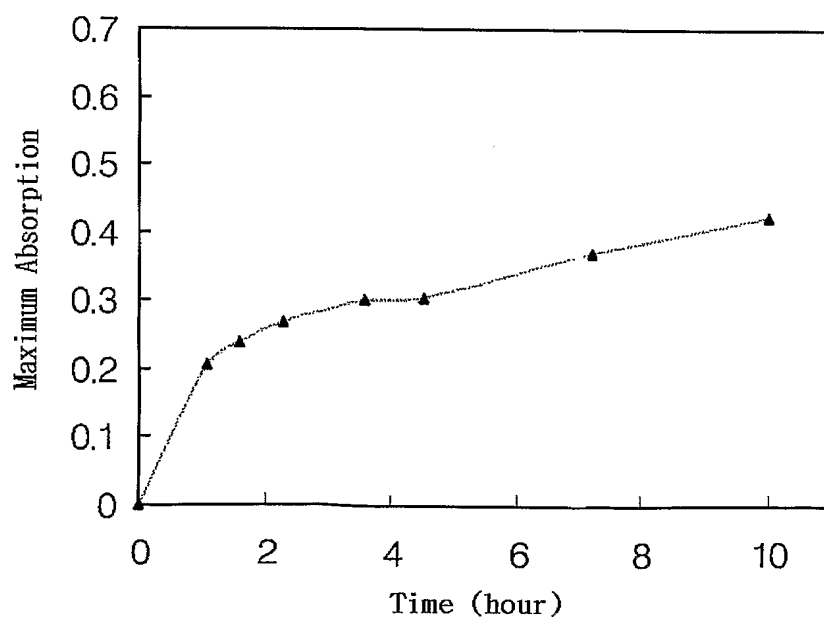


Fig. 24



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Fig. 25**Fig. 26**

INTERNATIONAL SEARCH REPORT

International application No.
PCT/KR02/01410

A. CLASSIFICATION OF SUBJECT MATTER**IPC7 A61K 7/00, A61K 7/48**

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC7: A61K ; C07D

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Korean Patents and applications for inventions since 1975

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

CA on CD; KIPASS; ESPACENET; Delphion

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	EP 1023897 A (ShOWA DENKO), 02. 08. 2000, p.3, line 14 - p.4, line 50	1-18
A	JP 11302145 A (HAMARI CHEM LTD), 02. 11. 1999, see the whole document	1-18
A	JP 03109478 A (KAO CORP), 09. 05. 1991, see the whole document	1-18
A	JP 63199281 A (SHISEIDO CO LTD), 17. 08. 1988, see the whole document	1-18
A	EP 38246 A (FABRE SA PIERRE), 21. 10. 1981, abstract	1-18
A	JP 55064511 A (KANEBO LTD), 15. 05. 1980, see the whole document	1-18
A	KR 2000-67126 A (PACIFIC CO LTD), 15.11.2000, see the whole document	1-18

☐ Further documents are listed in the continuation of Box C.

☒ See patent family annex.

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Date of the actual completion of the international search

11 NOVEMBER 2002 (11.11.2002)

Date of mailing of the international search report

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INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No.

PCT/KR02/01410

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JP 55064511 A	15. 05. 1980	none	
KR 2000-67126	15. 11. 2000	none	